

Arrhenius Equation

$$k = A \cdot e^{-\frac{E_a}{RT}}$$

$$R = 8.314 \frac{\text{J}}{\text{K mol}}$$

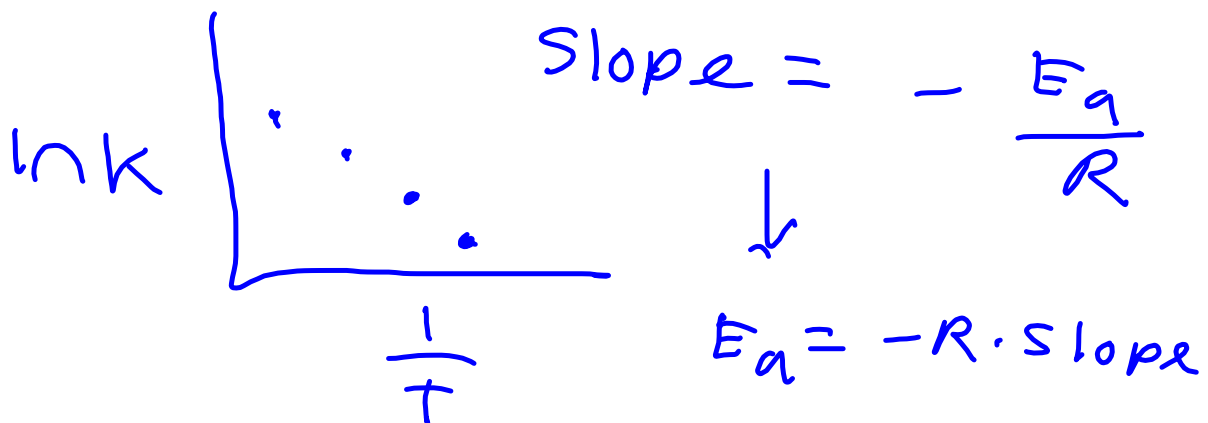
$$\ln k = \ln \left(A \cdot e^{-\frac{E_a}{RT}} \right)$$

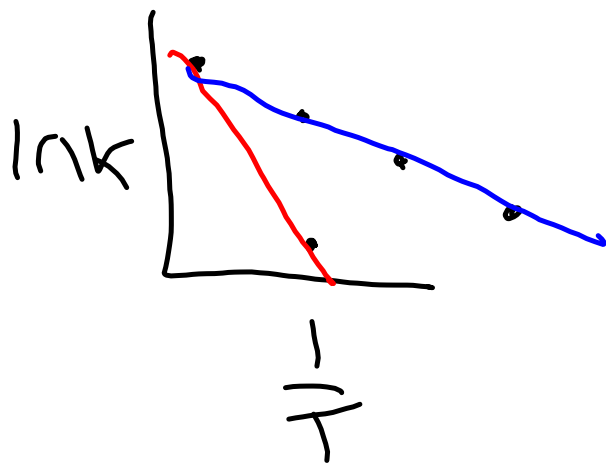
$$\ln k = \ln A + \cancel{\ln} \cancel{e}^{-\frac{E_a}{RT}}$$

$$\ln k = \ln A - \frac{E_a}{RT}$$

$$\ln k = \ln A - \frac{E_a}{R} \left(\frac{1}{T} \right)$$

$$y = b + m \cdot x$$





the 2-point Arrhenius equation

$$\ln k = \ln A - \left(\frac{E_a}{R}\right)\left(\frac{1}{T}\right)$$

$$(T_1, k_1) \quad (T_2, k_2)$$

$$\ln k_2 = \ln A - \left(\frac{E_a}{R}\right)\left(\frac{1}{T_2}\right)$$

$$- \left[\ln k_1 = \ln A - \left(\frac{E_a}{R}\right)\left(\frac{1}{T_1}\right) \right]$$

$$\ln k_2 - \ln k_1 = \left(\frac{E_a}{R}\right)\left(\frac{1}{T_1}\right) - \left(\frac{E_a}{R}\right)\left(\frac{1}{T_2}\right)$$

$$\ln\left(\frac{k_2}{k_1}\right) = \left(\frac{E_a}{R}\right)\left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

2-point Arrhenius Equation

Remember Clausius-Clapeyron?

$$\ln\left(\frac{P_2}{P_1}\right) = \frac{\Delta H_{\text{vap}}}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$ab - ac = a(b - c)$$



$$R = k[\text{NO}_2][\text{F}_2]$$

