

# Dalton's Law of Partial Pressure

$$P_{\text{TOTAL}} = P_1 + P_2 + \dots$$

$$PV = nRT$$

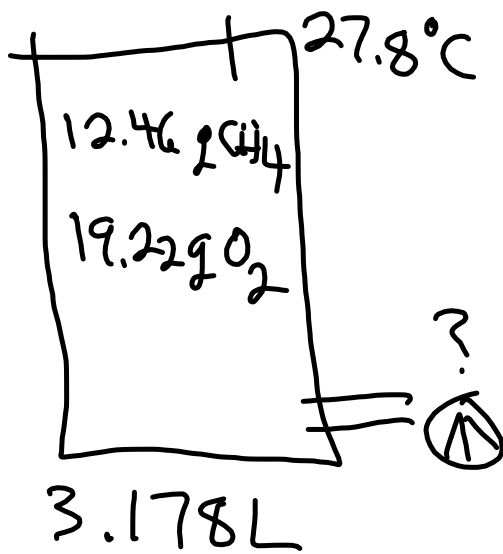
$$P_1 V = n_1 RT$$

$$P_2 V = n_2 RT$$

⋮

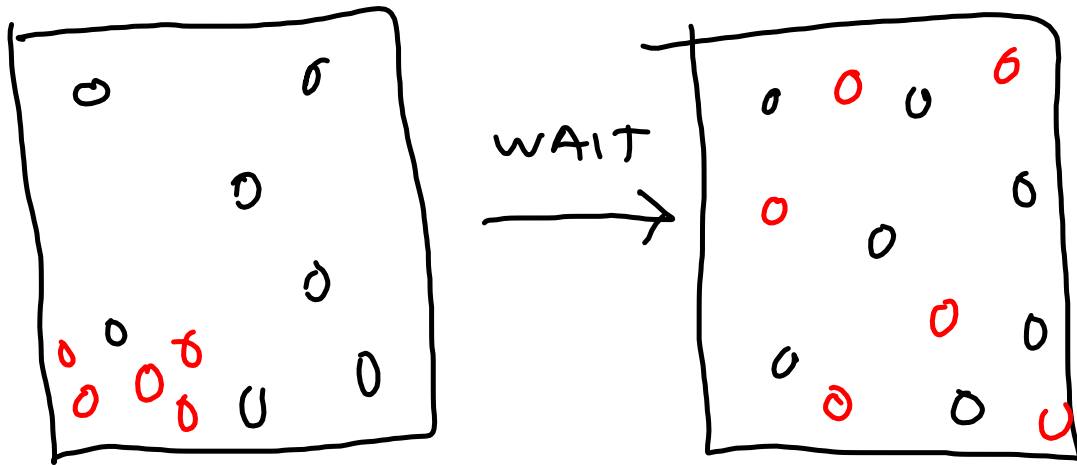
$$P_{\text{TOTAL}} V = n_{\text{TOTAL}} RT$$

$$n_{\text{TOTAL}} = n_1 + n_2 + \dots$$

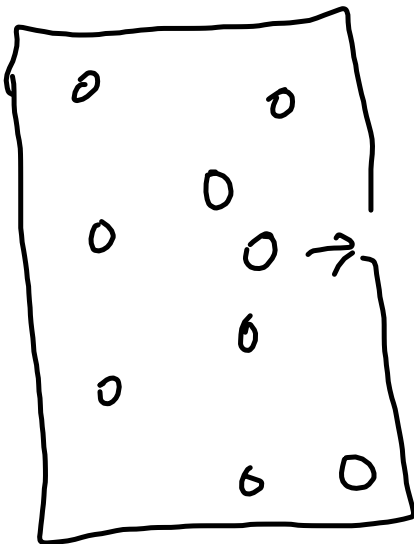


What total pressure is exerted by a gas mixture containing 12.46 g CH<sub>4</sub> and 19.22 g O<sub>2</sub> in a 3.178 L container at 27.8°C?

A steel tank having a volume of 9.133 L contains 7.515 g  $O_2$  and an unknown mass of  $N_2$  at  $23.6^\circ C$ . The total pressure exerted by this gas mixture is 9.975 atm. What mass of  $N_2$  is in the tank?



DIFFUSION



VACUUM

$H_2$  2.016

$O_2$  32.00

$$\overline{KE} \propto T$$

$$\overline{KE} = \frac{1}{2} m \overline{v^2}$$

# Graham's Law of Effusion

$$E.R. \propto \frac{1}{\sqrt{MWT}}$$

$$E.T. \propto \sqrt{MWT}$$

$$E.T. = k \sqrt{MWT}$$

$$k = f(P, T, A)$$

$$\frac{E.T._1}{E.T._2} = \frac{\cancel{k} \sqrt{MWT_1}}{\cancel{k} \sqrt{MWT_2}}$$

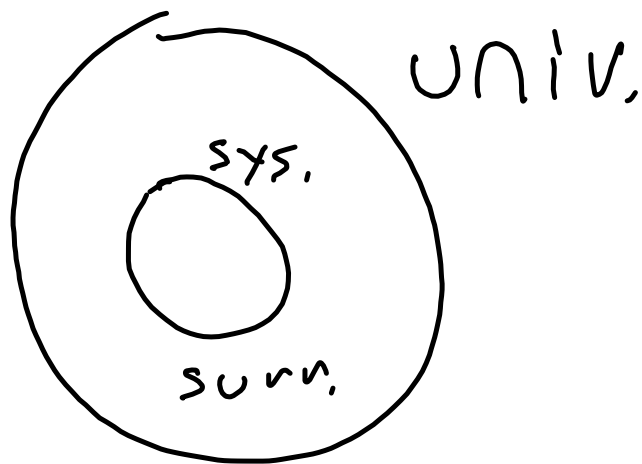
$$\frac{E.T._1}{E.T._2} = \frac{\sqrt{MWT_1}}{\sqrt{MWT_2}} = \sqrt{\frac{MWT_1}{MWT_2}}$$

$$\frac{E.R._1}{E.R._2} = \sqrt{\frac{MWT_2}{MWT_1}}$$

If it takes 37.15 min for  $\text{CO}_2$  to effuse through a small hole in its container, how long would it take for  $\text{NH}_3$  to effuse under the same conditions of temp and pressure? (comparisons being for the same number of moles of effused gas.)

An unknown gas requires 60.0 minutes to effuse through a small hole in its container. Under the same conditions of temp. and pressure, helium (He) requires 30.0 min. What is the molecular weight of the unknown gas?

# THERMOCHEMISTRY



Open system - can exchange both matter and energy with the surroundings

Closed system - can exchange energy, but not matter, with the surroundings.

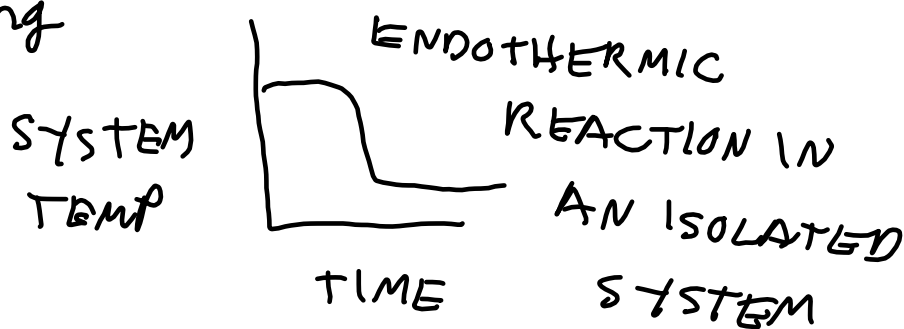
Isolated system - Exchanges neither matter nor energy with the surroundings.



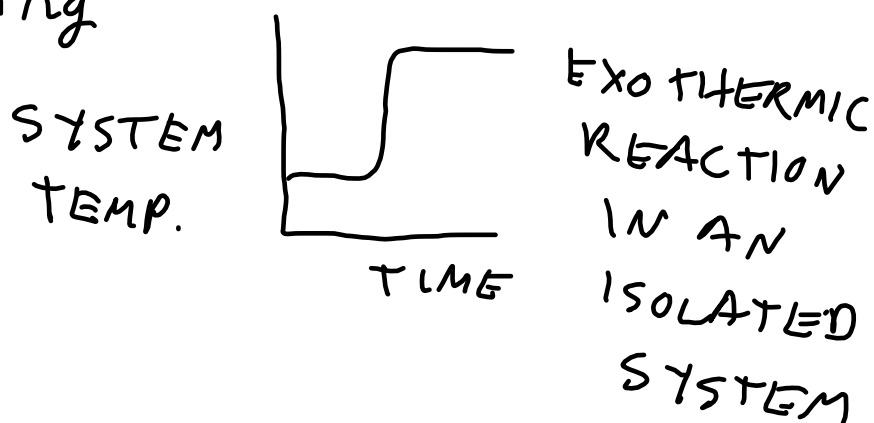


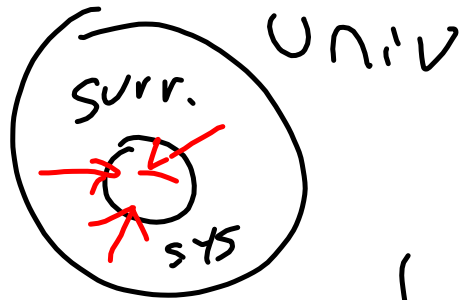
Chemical reactions can be:  
 Endothermic - absorb heat  
 or  
 Exothermic - release heat  
 or  
 Neither (rare)

Endothermic reactions cause cooling

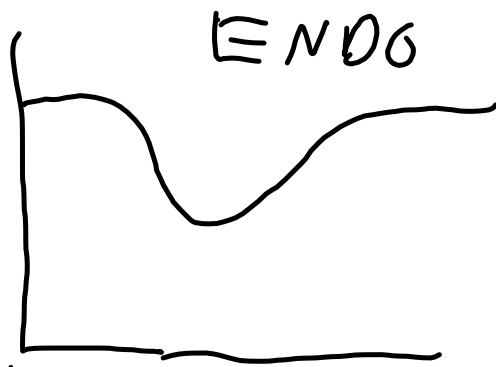


Exothermic reactions cause warming





→ SYS.  
 TEM



closed (not isolated) TIME  
 system

↘ SYS.  
 TEMP

