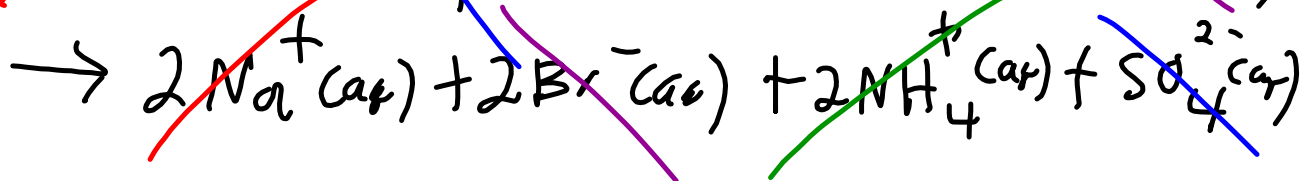
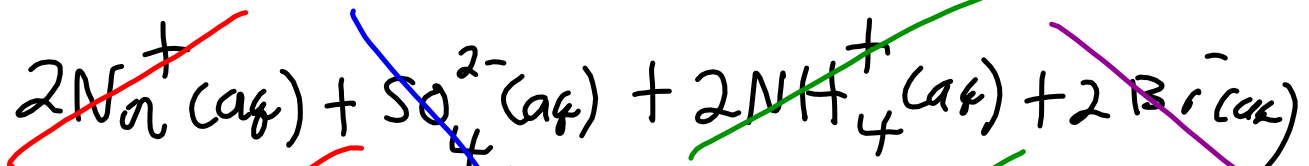
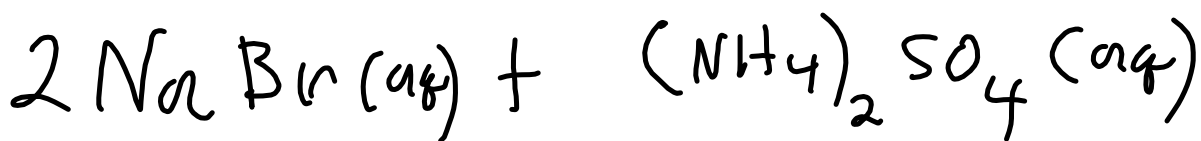
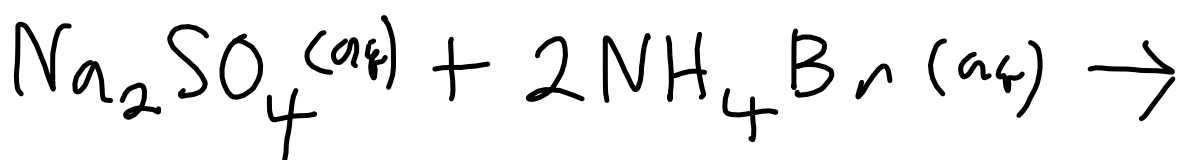
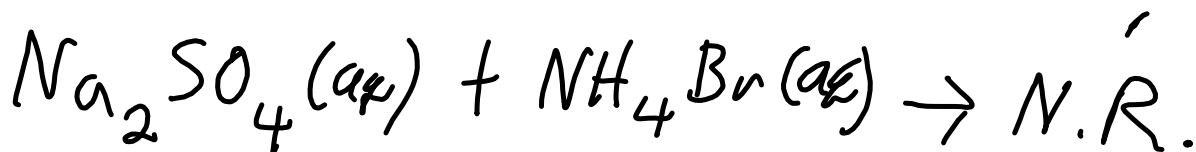


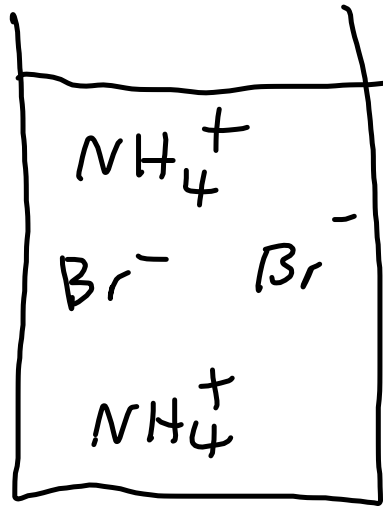
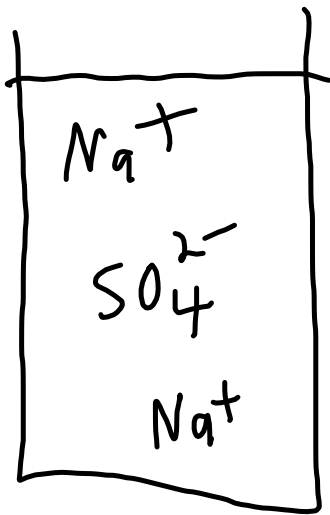
sodium sulfate + ammonium bromide →

sodium bromide + ammonium sulfate

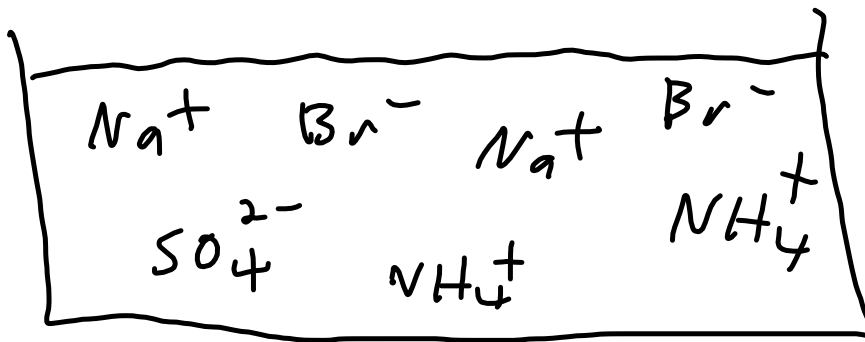


NO NET IONIC EQN |





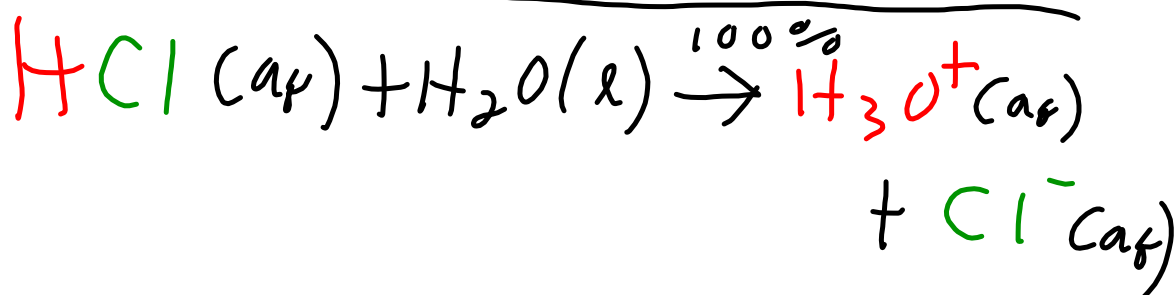
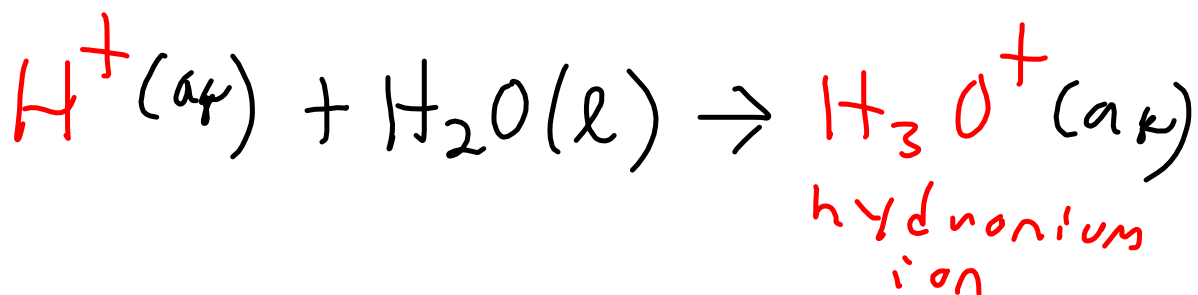
Mix



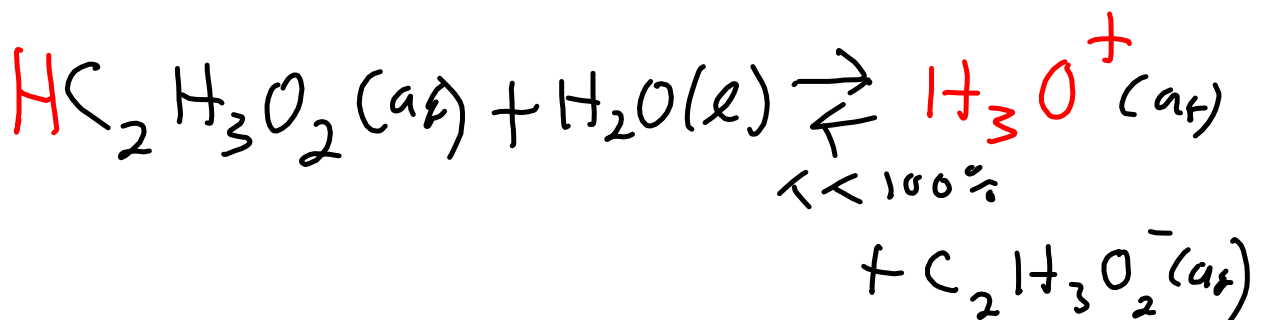
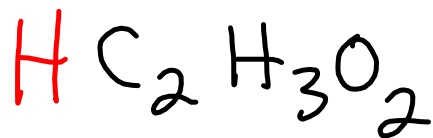
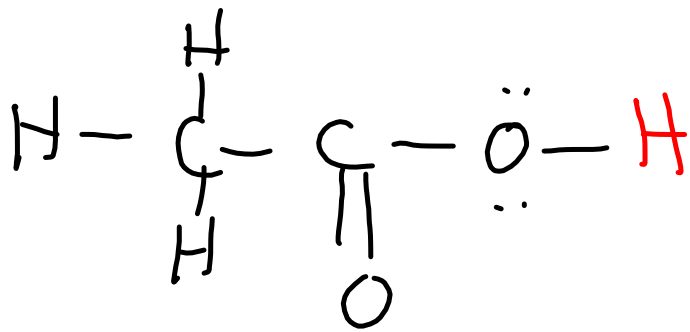
STABLE

# ACID / BASE CHEMISTRY

An **ACID** is a substance that produces  $H^+$  (or equivalently)  $H_3O^+$  cations when dissolved in water

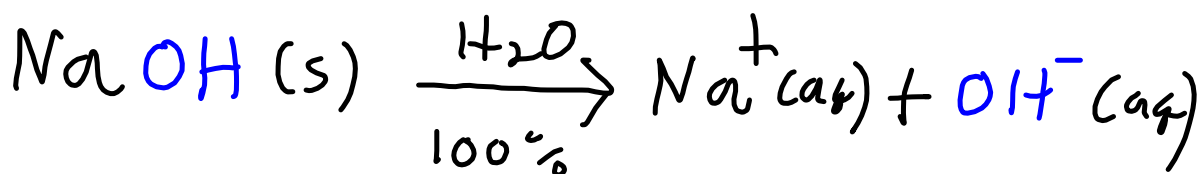


acetic acid

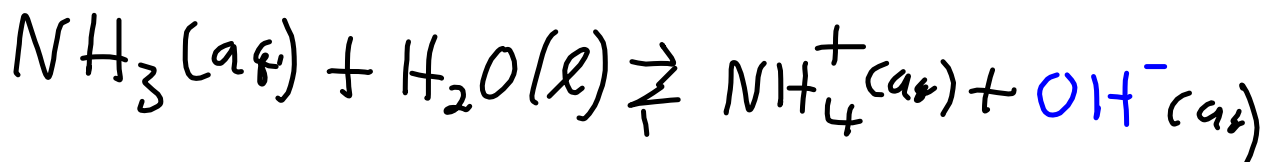


A **BASE** is a substance that produces  $\text{OH}^-$  ions when dissolved in water.

**STRONG BASES** are usually water soluble hydroxide salts.



Weak **bases** are usually molecular compounds which don't have the OH group explicitly in their formula, but which react with water to produce  $\text{OH}^-$  ions.



|      |        |                 |      |
|------|--------|-----------------|------|
|      |        | BASE            |      |
|      |        | STRONG          | weak |
| ACID | STRONG | NEUTRAL<br>SALT |      |
|      | weak   |                 |      |

