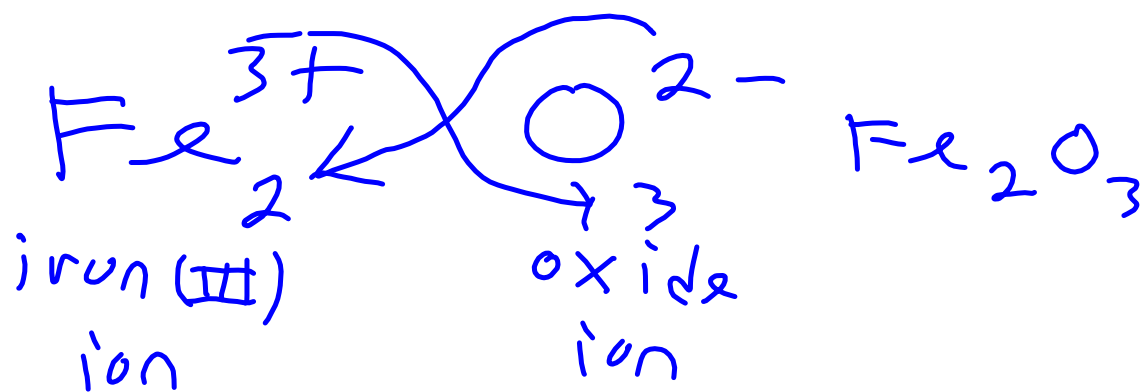


Stock NAME	-ic/-ous NAME	FORMULA
iron(III) oxide	ferric oxide	$Fe_2O_3$
copper(II) chloride	cupric chloride	$CuCl_2$
lead(II) sulfide	plumbous sulfide	$PbS$
manganese(II) nitride	manganous nitride	$Mn_3N_2$
cobalt(II) bromide	cobaltous bromide	$CoBr_2$

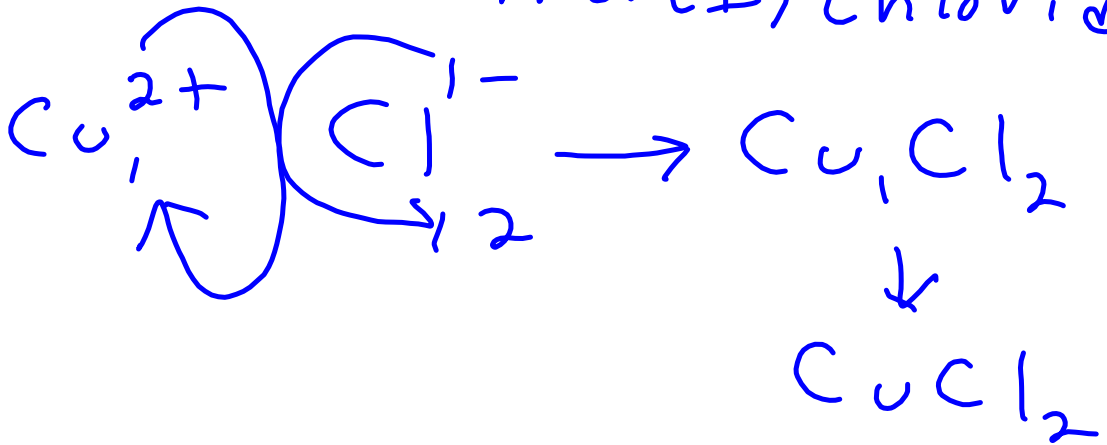


$Fe^{3+}$  ferric ion

$Fe^{2+}$  ferrous ion

$\text{Cu}^{2+}$  cupric ion  
copper(II) ion

$\text{Cu}^+$  cuprous ion  
copper(I) chloride



EACH  $\text{Pb}^{2+}$   $\text{S}^{2-}$

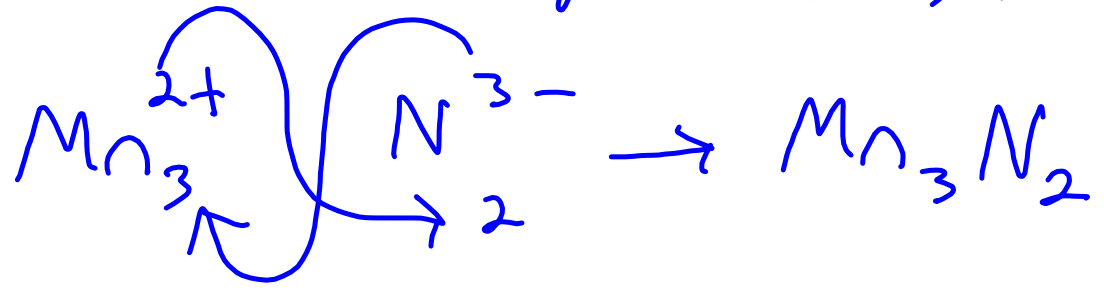
TOTAL  $2+$   $2-$  =  $\emptyset$

$\text{Pb}^{2+}$  lead(II) ion  
Plumbous ion

$\text{Pb}^{4+}$  lead(IV) ion  
Plumbic ion

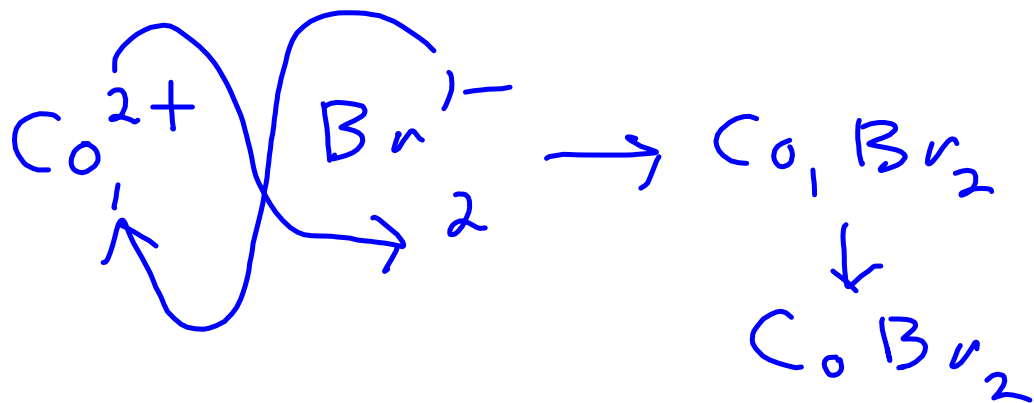
$Mn^{3+}$  manganic ion  
manganese(III) ion

$Mn^{2+}$  manganous ion  
manganese(II) ion



$\text{Co}^{3+}$  cobalt(III) ion  
cobaltic ion

$\text{Co}^{2+}$  cobalt(II) ion  
cobaltous ion



## POLYATOMIC IONS

$\text{NH}_4^+$  ammonium ion

$\text{H}_3\text{O}^+$  hydronium ion  
(only in aqueous solution)

$\text{SO}_4^{2-}$  sulfate ion

$\text{SO}_3^{2-}$  sulfite ion

$\text{S}^{2-}$   
sulfide  
ion

$\text{NO}_3^-$  nitrate

$\text{NO}_2^-$  nitrite

$\text{N}^{3-}$   
nitride  
ion

$\text{PO}_4^{3-}$  phosphate  
ion

$\text{PO}_3^{3-}$  phosphite  
ion

$\text{P}^{3-}$  phosphide  
ion



$\text{ClO}_4^-$  perchlorate ion

$\text{ClO}_3^-$  chlorate ion

$\text{ClO}_2^-$  chlorite ion

$\text{ClO}^-$  hypochlorite ion

$\text{Cl}^-$  chloride ion

$\text{CO}_3^{2-}$  carbonate  
ion

$\text{OH}^-$  hydroxide  
ion

$\text{CN}^-$  cyanide  
ion

$\text{O}_2^{2-}$  Peroxide  
ion

Write formulas for the following compounds containing polyatomic ions.

1. sodium carbonate

2. ammonium chloride

3. calcium phosphate

4. magnesium nitrite

