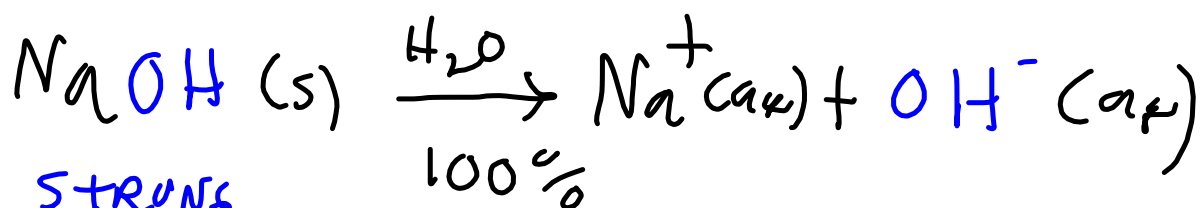


STRONG  
ACID



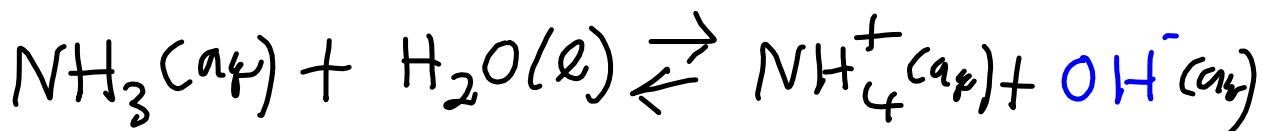
**STRONG BASES** are usually water-soluble hydroxide salts

**Weak bases** are usually molecular compounds which do not have the OH group as part of their formula, but which react with  $H_2O$  to form  $OH^-$  ions as a product of the reaction.



**STRONG  
BASE**

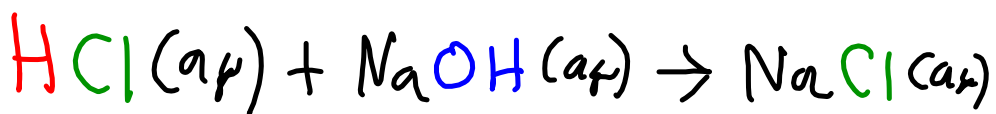
$\ll 100\%$



**weak  
base**

# NEUTRALIZATION REACTIONS

A C D	STRONG	neutral salt	acidic salt
	weak	basic salt	
		STRONG	weak
		BASE	



STRONG  
ACID

STRONG  
BASE

+  $\text{H}_2\text{O}(\text{l})$

