

A comparison of molecular and ionic compounds dissolving in water.

Molecular Example:

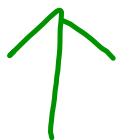


Ionic compounds:

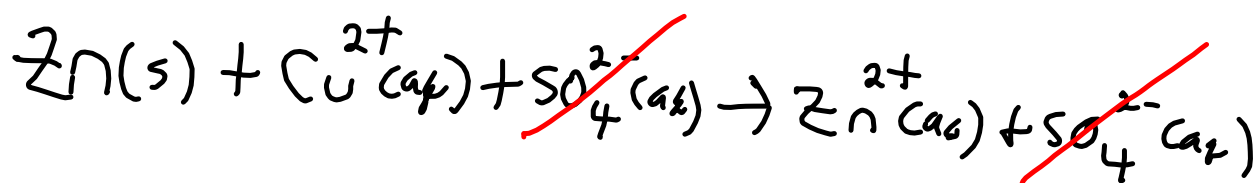
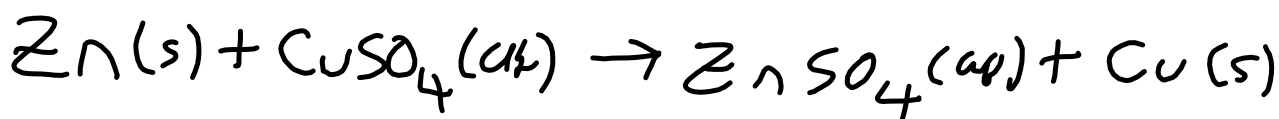
You might be tempted to write something like this:



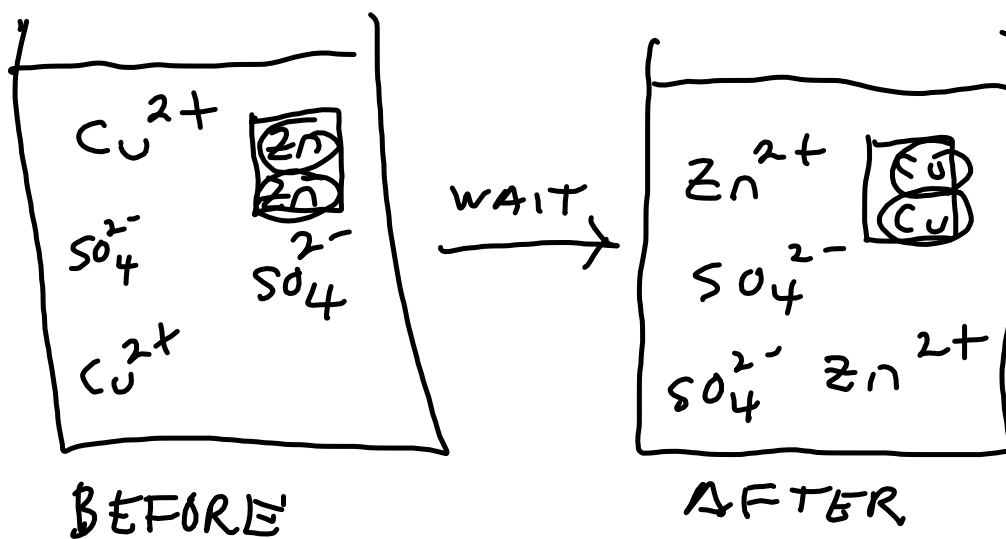
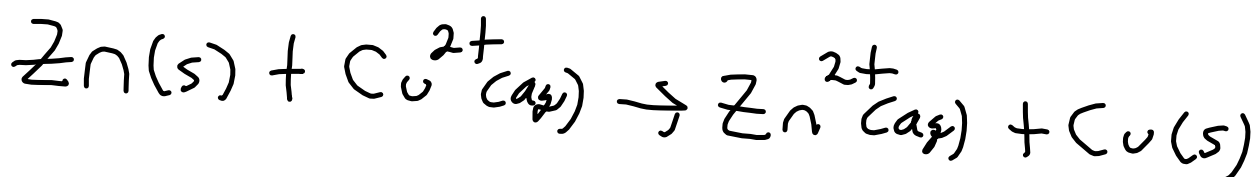
NOT what really happens!

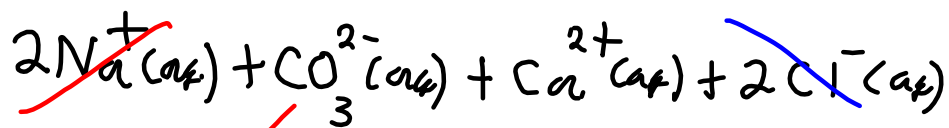
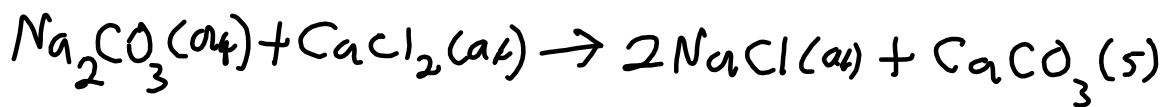


What really happens

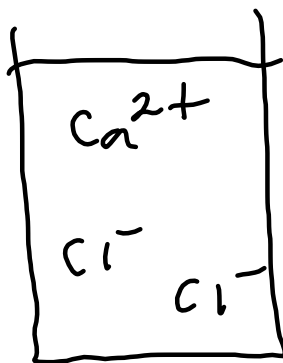
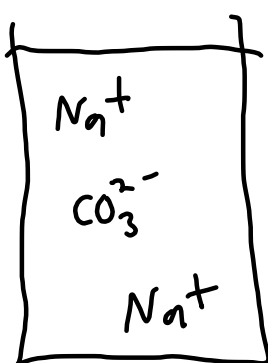
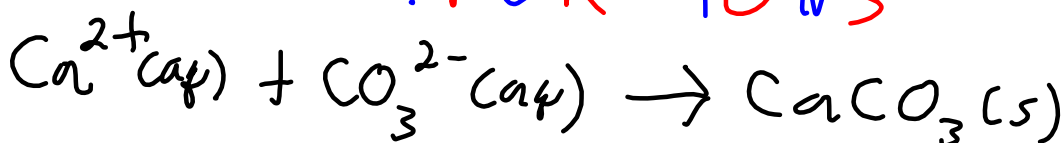


SPECTATOR ION + Cu(s)

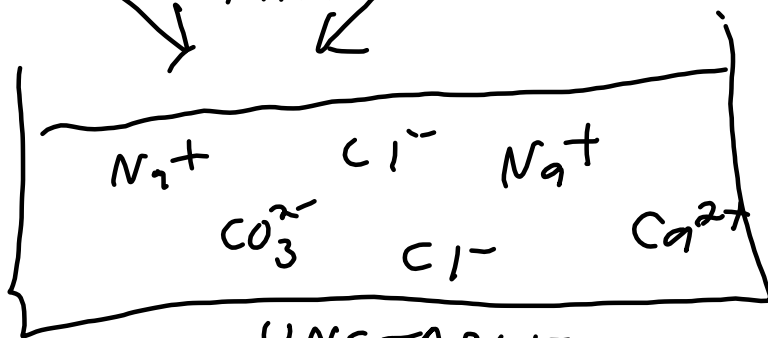




SPECTATOR IONS

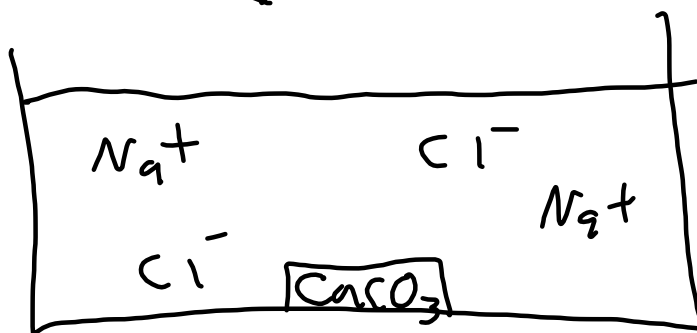


MIX



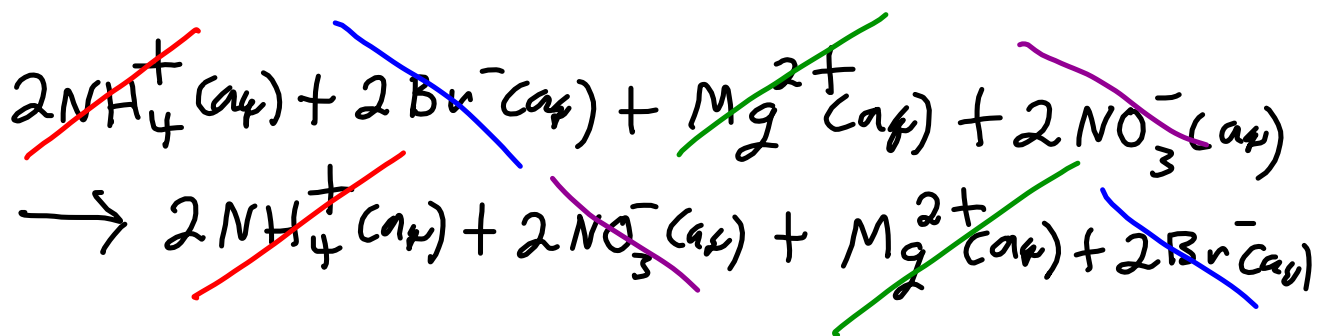
UNSTABLE

↓

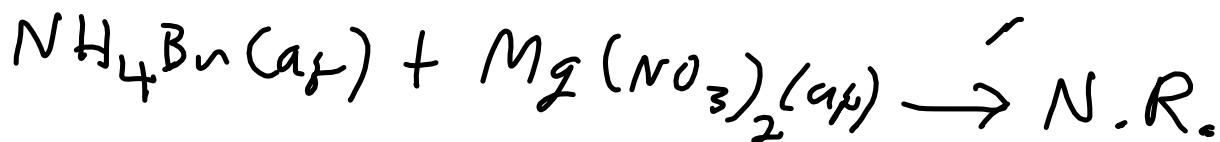


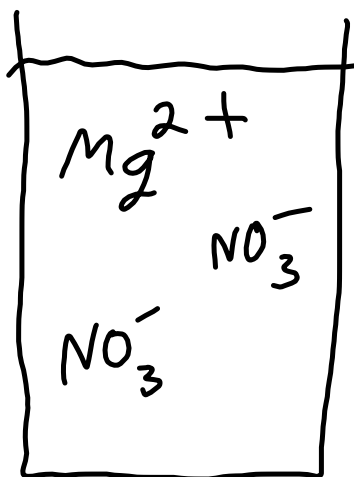
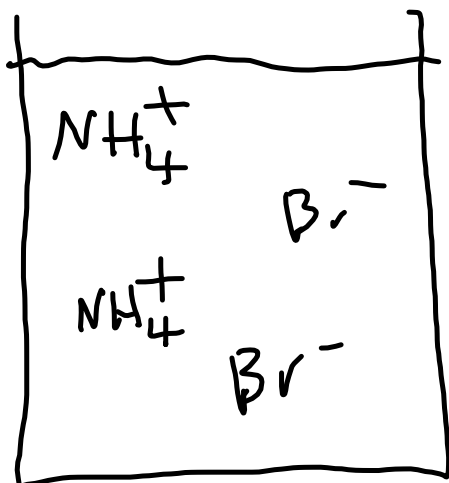
ammonium bromide + magnesium nitrate \rightarrow

ammonium nitrate + magnesium bromide

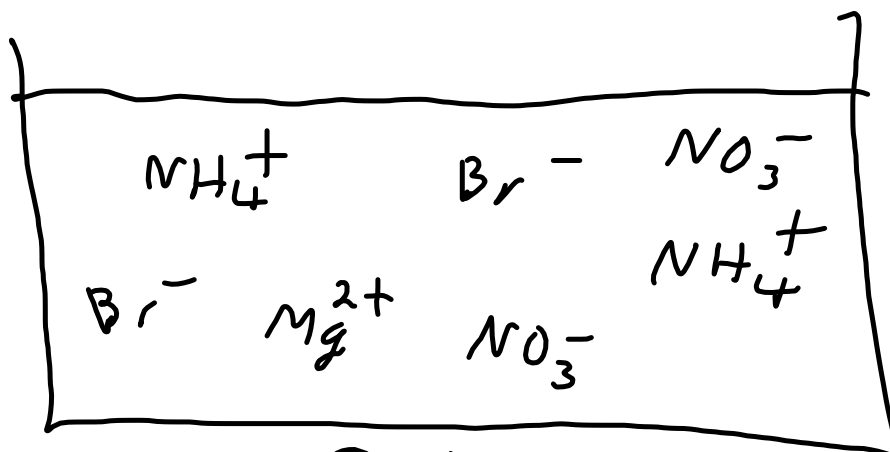


NO NET IONIC EQN. |





MIX



STABLE

ACID - A substance that produces H^+ ion, or equivalently, H_3O^+ ion, when dissolved in H_2O

BASE - A substance that produces OH^- ions when dissolved in H_2O .

Example of an ACID

