

CaCl_2 calcium chloride
~~calcium dichloride~~

NO

NO_2

N_2O

N_2O_3

N_2O_4

N_2O_5

1 -	mono -	8	octa -
2 -	di -	9	nona -
3 -	tri -	10	deca -
4 -	tetra -		
5	penta -		
6	hexa -		
7	hepta -		

NO nitrogen monoxide

NO_2 nitrogen dioxide

N_2O dinitrogen monoxide

N_2O_3 dinitrogen trioxide

N_2O_4 dinitrogen tetroxide

N_2O_5 dinitrogen pentoxide

$PbCl_4$ ~~lead tetrachloride~~
lead(~~IV~~) chloride
~~plumbic chloride~~

$PbCl_2$ ~~lead dichloride~~
lead(II) chloride
plumbous chloride

ACIDS

1. Binary
2. Ternary

Binary Acids

Composed of hydrogen and one other element, usually a non-metal. Binary acids are only named as acids when dissolved in H_2O

hydro _____ ic acid

HF hydrofluoric acid

HCl hydrochloric acid

HBr hydrobromic acid

HI hydroiodic acid

H_2S hydrosulfuric acid

Ternary acids
Hydrogen, oxygen, and
one other element.

Note: We can recognize
polyatomic anions in their
formulas

SO_4^{2-} sulfate
ion

H_2SO_4
sulfuric acid

SO_3^{2-} sulfite
ion

H_2SO_3
sulfurous acid

ate
ion



ic
acid

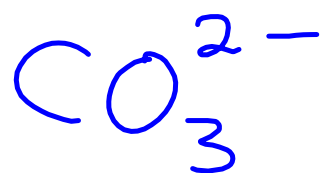
ite
ion



ous
acid

NAME	FORMULA
Carbonic acid	H_2CO_3
nitrous acid	HNO_2
Phosphoric acid	H_3PO_4
nitric acid	HNO_3
hydrochloric acid	HCl

Carbonate ion



H_2CO_3 carbonic acid

HNO_2 nitrous acid

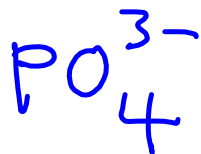


Phosphoric acid

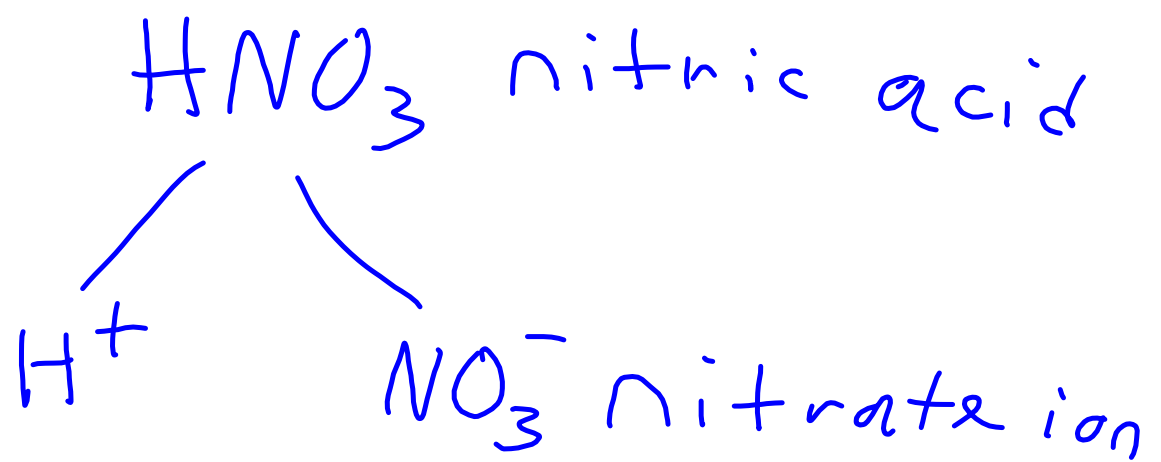
Phosphate ion

↓ minor spelling
change

phosphate ion



H_3PO_4 phosphoric acid



ACID ANIONS & ACID SALTS

SO_4^{2-}
sulfate
ion

HSO_4^-
hydrogen
sulfate
ion

H_2SO_4
sulfuric
acid

Alternate

Name:

bisulfate ion

PO_4^{3-} phosphate ion

HPO_4^{2-} monohydrogen phosphate ion
biphosphate ion

H_2PO_4^- dihydrogen phosphate ion

H_3PO_4 phosphoric acid