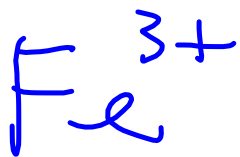
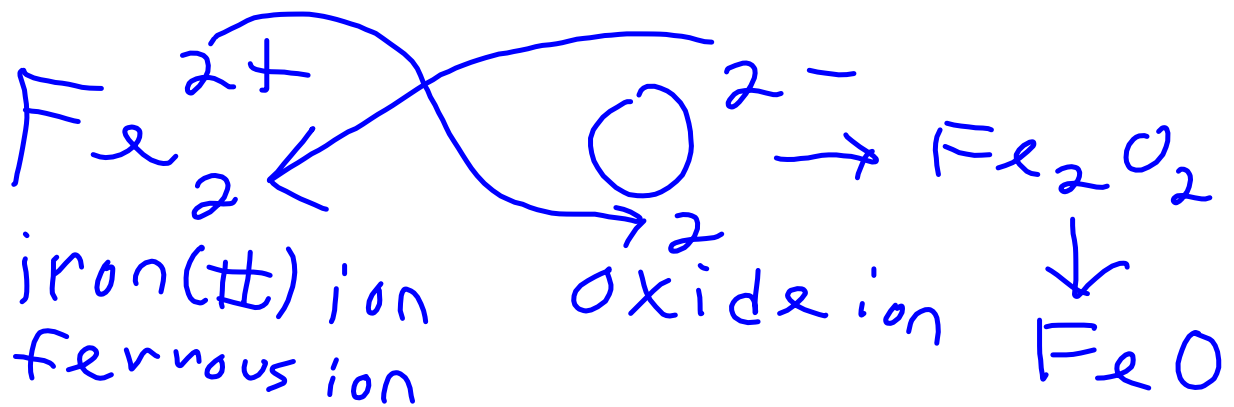
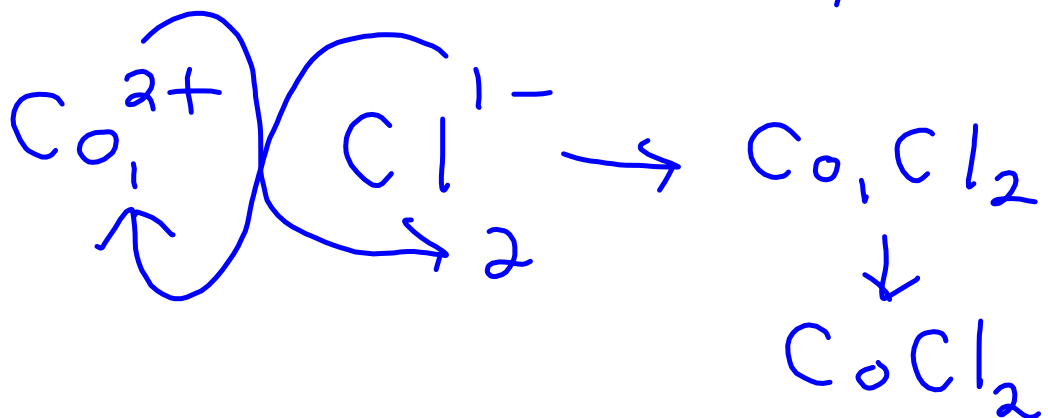


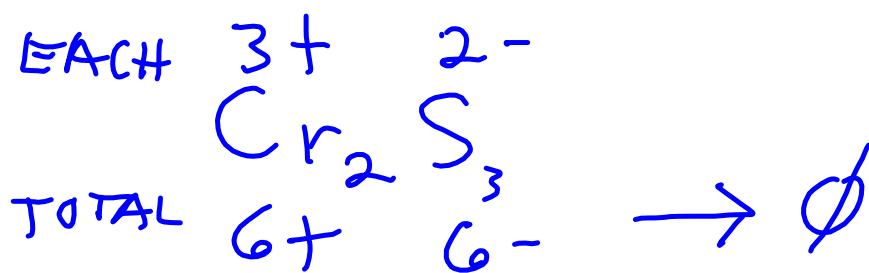
STOCK NAME	-ic/-ous NAME	FORMULA
iron(II) oxide	ferrous oxide	FeO
cobalt(II) chloride	cobaltous chloride	CoCl_2
chromium(III) sulfide	chromic sulfide	Cr_2S_3
lead(IV) oxide	plumbic oxide	PbO_2
manganese(III) nitride	manganic nitride	MnN



iron(III) ion
 ferric ion

Co^{3+} cobaltic ion
cobalt(III) ion
 Co^{2+} cobaltous ion
cobalt(II) ion





Cr^{3+} Chromium(III) ion
chromic ion

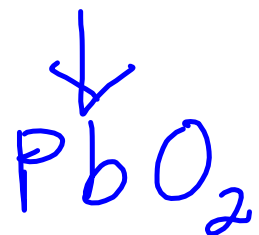
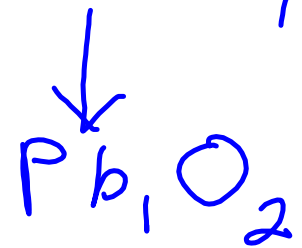
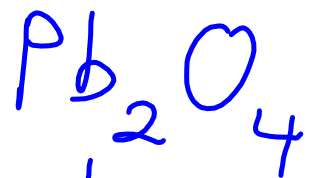
Cr^{2+} Chromium(II) ion
chromous ion

Lead (IV) oxide



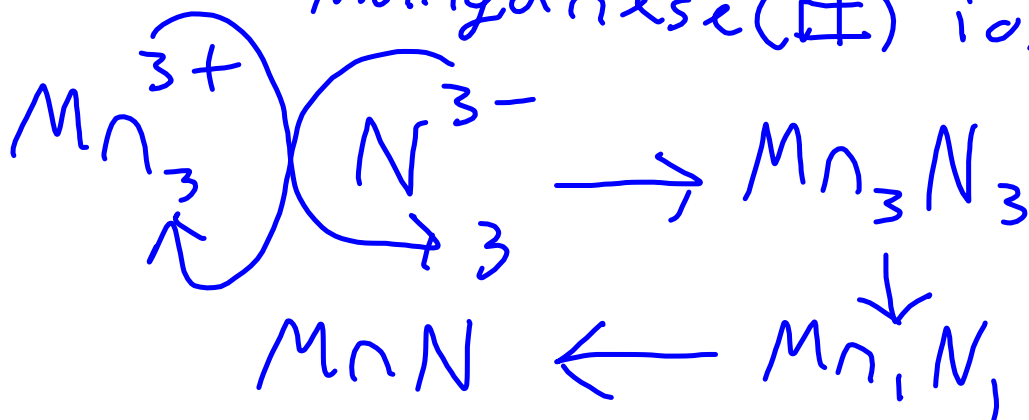
Pb^{4+} plumbic ion

Pb^{2+} plumbous ion



Manganic nitride

Mn^{3+} manganic ion
 manganese(III) ion
 Mn^{2+} manganous ion
 manganese(II) ion



POLYATOMIC IONS

NH_4^+ ammonium

H_3O^+ hydronium ion
(only in aqueous solution)

SO_4^{2-} sulfate ion

SO_3^{2-} sulfite ion

-ate = more -ite = less

S^{2-} sulfide ion

PO_4^{3-} phosphate ion

PO_3^{3-} phosphite ion

P^{3-} phosphide ion

ClO_4^- perchlorate

ClO_3^- chlorate ion

ClO_2^- chlorite ion

ClO^- hypochlorite ion

Cl^- chloride ion

NO_3^- nitrate ion

NO_2^- nitrite ion

N^{3-} nitride ion

CO_3^{2-} carbonate ion

OH^- hydroxide ion

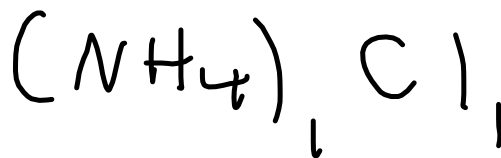
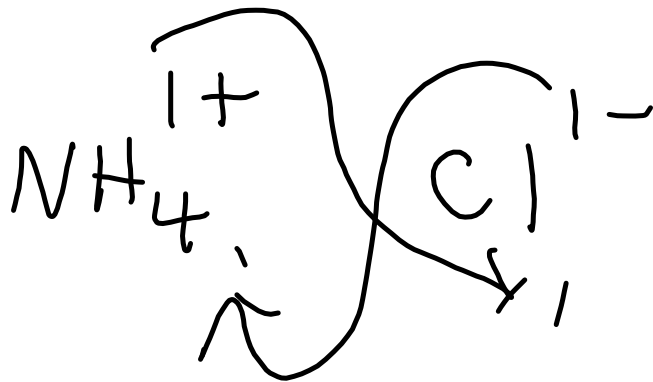
CN^- cyanide ion

O_2^{2-} peroxide ion

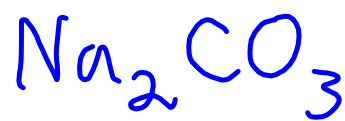
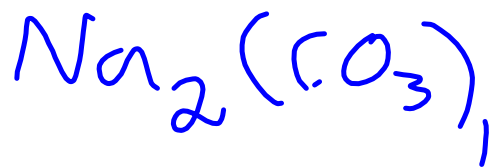
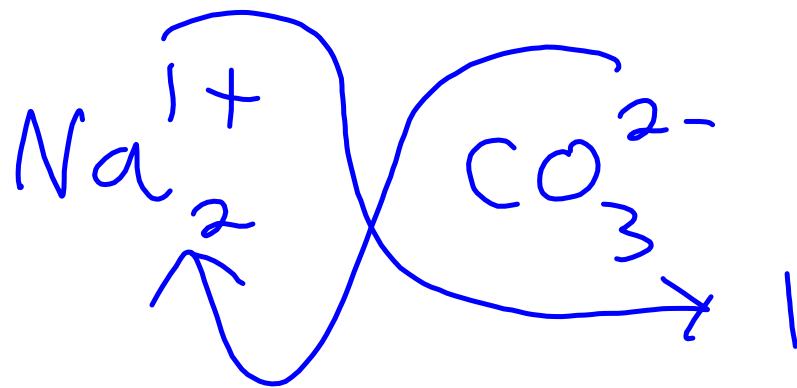
Write formulas for the following compounds containing polyatomic ions.

1. ammonium chloride NH_4Cl
2. sodium carbonate Na_2CO_3
3. barium nitrite $\text{Ba}(\text{NO}_2)_2$
4. calcium phosphate $\text{Ca}_3(\text{PO}_4)_2$

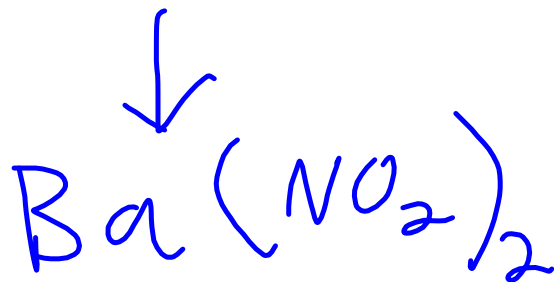
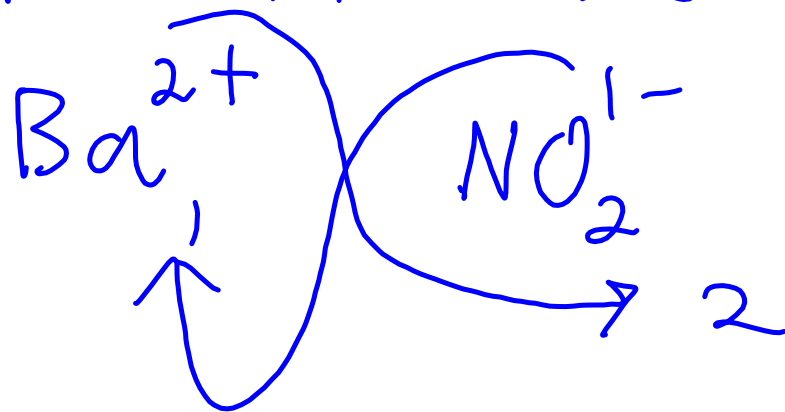
ammonium chloride



Sodium carbonate



barium nitrite



Calcium phosphate

